



Pearson
Edexcel

Mark Scheme (Results)

November 2020

Pearson Edexcel International GCSE
In Chemistry (4CH1) Paper 1CR and Science
(Double Award) (4SD0) Paper 1CR

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General Marking Guidance

- All candidates must receive the same treatment. Examiners must mark the first candidate in exactly the same way as they mark the last.
- Mark schemes should be applied positively. Candidates must be rewarded for what they have shown they can do rather than penalised for omissions.
- Examiners should mark according to the mark scheme not according to their perception of where the grade boundaries may lie.
- There is no ceiling on achievement. All marks on the mark scheme should be used appropriately.
- All the marks on the mark scheme are designed to be awarded. Examiners should always award full marks if deserved, i.e. if the answer matches the mark scheme. Examiners should also be prepared to award zero marks if the candidate's response is not worthy of credit according to the mark scheme.
- Where some judgement is required, mark schemes will provide the principles by which marks will be awarded and exemplification may be limited.
- When examiners are in doubt regarding the application of the mark scheme to a candidate's response, the team leader must be consulted.
- Crossed out work should be marked UNLESS the candidate has replaced it with an alternative response.

Question number	Answer	Notes	Marks
1 (a) (i)	simple distillation	REJECT fractional distillation REJECT distillation	1
(ii)	chromatography		1
(iii)	fractional distillation		1
(b)	M1 dissolve M2 solute M3 solvent		3
			Total 6

Question number	Answer	Notes	Marks
2 (a) (i)	Na/K	ACCEPT sodium/potassium	1
(ii)	He	ACCEPT helium	1
(iii)	Br	ACCEPT bromine REJECT bromide	1
(iv)	Na and Cl	ACCEPT sodium and chlorine REJECT chloride	1
(b)	2.8.1	ACCEPT 2,8,1 or 2 8 1 ALLOW diagram of atom showing correct electron configuration.	1
			Total 5

Question number	Answer	Notes	Marks										
3 (a)	<table border="1"> <tr> <td>molecular formula</td> <td>C_3H_6</td> </tr> <tr> <td>name of this alkene</td> <td>propene</td> </tr> <tr> <td>empirical formula</td> <td>CH_2</td> </tr> <tr> <td>general formula</td> <td>C_nH_{2n}</td> </tr> <tr> <td>displayed formula</td> <td> <pre> H H H H - C = C - C - H H </pre> </td> </tr> </table>	molecular formula	C_3H_6	name of this alkene	propene	empirical formula	CH_2	general formula	C_nH_{2n}	displayed formula	<pre> H H H H - C = C - C - H H </pre>	<p>ACCEPT propylene</p> <p>ACCEPT N or other letters e.g. x</p>	4
molecular formula	C_3H_6												
name of this alkene	propene												
empirical formula	CH_2												
general formula	C_nH_{2n}												
displayed formula	<pre> H H H H - C = C - C - H H </pre>												
	1 mark for each correct answer												
b) i)	(contains a carbon carbon) double bond	ACCEPT multiple bond	1										
ii)	M1 add bromine water	REJECT bromine REJECT bromide	2										
	M2 decolourises/changes to colourless	ALLOW turns colourless IGNORE clear M2 dep M1 or near miss											
(c) (i)	$CH_4 + Cl_2 \rightarrow CH_3Cl + HCl$	ACCEPT multiples	1										
(ii)	D Substitution		1										
	A is incorrect as it is not an addition reaction B is incorrect as it is not a decomposition reaction C is incorrect as it is not a neutralisation reaction												
(iii)	Ultraviolet radiation/light	ACCEPT UV/ultraviolet rays	1										
(d) (i)	M1 (isomers have) the same molecular formula		2										
	M2 (but) different structural/displayed formulae	ALLOW different structures ALLOW different arrangement of atoms											

(ii)	M1	$ \begin{array}{c} \text{H} \quad \text{H} \\ \quad \\ \text{H}-\text{C}-\text{C}-\text{H} \\ \quad \\ \text{Cl} \quad \text{Cl} \end{array} $	2
	M2	$ \begin{array}{c} \text{Cl} \quad \text{H} \\ \quad \\ \text{H}-\text{C}-\text{C}-\text{H} \\ \quad \\ \text{Cl} \quad \text{H} \end{array} $	
			Total 14

Question number	Answer	Notes	Marks
4 (a)	$2 \text{H}_2\text{O}_2 \rightarrow 2 \text{H}_2\text{O} + (1) \text{O}_2$	ALLOW multiples and fractions	1
(b)	Relights a glowing splint/spill		1
(c)	Speeds up/increases rate of the reaction	IGNORE references to lowering activation energy	1
(d) (i)	All points plotted correctly	ALLOW \pm half a square	1
(ii)	Point at 8 minutes circled		1
(iii)	Smooth curve of best fit		1
(iv)	took the reading too soon/before 8 minutes	ACCEPT misread the volume (of oxygen)	1
(v)	M1 vertical line on graph drawn to curve from 3 mins M2 value obtained from candidate's graph	Expected value 29 or 30cm^3 ALLOW \pm half a square	2

Question number	Answer	Notes	Marks
4 (e) (i)	<p>M1 curve drawn on graph that is less steep than curve of student's results.</p> <p>M2 curve levels off at 40 cm³</p>	<p>ALLOW ± half a square</p>	2
	<p>(ii) An explanation that links the following three points</p> <p>M1 reaction is slower</p> <p>M2 fewer particles/molecules (in the same volume)</p> <p>M3 fewer collisions per unit time</p>	<p>ACCEPT particles are further apart / less crowded</p> <p>ACCEPT less frequent collisions</p> <p>IGNORE less chance of a collision</p>	3
			Total 14

Question number	Answer	Notes	Marks
5 (a) (i)	6/six		1
(ii)	One of the following two points methanol/it) does not contain only carbon and hydrogen OR (methanol/it) contains (an atom of) oxygen		1
(b) (i)	M1 two/ pair of electrons M2 shared between two atoms	ACCEPT (electrons) attracted to the nuclei (of the two atoms in the bond) ACCEPT M1 (electrostatic) attraction between two nuclei M2 (and the) shared pair(s) of electrons (between them)	2
(ii)	M1 4 pairs of electrons around central carbon atom, with one pair to O and 3 pairs to H M2 rest of molecule fully correct	ALLOW any combination of dots and crosses M2 DEP on M1	2

Question number	Answer	Notes	Marks												
5 (c) (i)	<ul style="list-style-type: none"> • Divide percentages by relative atomic masses • Divide results by smallest value to obtain ratio • Write empirical formula <p>Example calculation</p> <p>M1</p> <table style="margin-left: 40px;"> <tr> <td>C</td> <td>H</td> <td>O</td> </tr> <tr> <td>$\frac{38.7}{12}$</td> <td>$\frac{9.7}{1}$</td> <td>$\frac{51.6}{16}$</td> </tr> </table> <p>M2</p> <table style="margin-left: 40px;"> <tr> <td>$\frac{3.225}{(3.225)}$</td> <td>$\frac{9.7}{(3.225)}$</td> <td>$\frac{3.225}{(3.225)}$</td> </tr> </table> <p>OR</p> <table style="margin-left: 40px;"> <tr> <td>1</td> <td>3</td> <td>1</td> </tr> </table> <p>M3</p> <p>CH₃O</p>	C	H	O	$\frac{38.7}{12}$	$\frac{9.7}{1}$	$\frac{51.6}{16}$	$\frac{3.225}{(3.225)}$	$\frac{9.7}{(3.225)}$	$\frac{3.225}{(3.225)}$	1	3	1	<p>0 marks if division by atomic numbers or upside-down calculation</p> <p>M2 subsumes M1</p> <p>ACCEPT symbols in any order</p>	3
C	H	O													
$\frac{38.7}{12}$	$\frac{9.7}{1}$	$\frac{51.6}{16}$													
$\frac{3.225}{(3.225)}$	$\frac{9.7}{(3.225)}$	$\frac{3.225}{(3.225)}$													
1	3	1													
(ii)	<ul style="list-style-type: none"> • Divide relative molecular mass by empirical formula mass • Write molecular formula <p>Example calculation</p> <p>M1</p> <table style="margin-left: 40px;"> <tr> <td>$\frac{62}{12 + 3 + 16}$</td> <td>OR</td> <td>$\frac{62}{31}$</td> <td>(= 2)</td> </tr> </table> <p>M2</p> <p>C₂H₆O₂</p>	$\frac{62}{12 + 3 + 16}$	OR	$\frac{62}{31}$	(= 2)	<p>ACCEPT symbols in any order</p> <p>Correct answer without working scores 2 marks</p>	2								
$\frac{62}{12 + 3 + 16}$	OR	$\frac{62}{31}$	(= 2)												
Total 11															

Question number	Answer	Notes	Marks
6 (a) (i)	halogens	REJECT halides	1
(ii)	(pale) green		1
(iii)	M1 test with (damp) litmus paper M2 bleaches	ALLOW (damp) universal indicator paper ALLOW turns white IGNORE turns red M2 dep on M1	2
(b)	A description that refers to the following three points M1 add (dilute) nitric acid (to the unknown solution) M2 add silver nitrate (solution) M3 (pale) yellow precipitate	M1 and M2 can be in either order REJECT addition of incorrect acid M3 dep on M2	3

Question number	Answer	Notes	Marks
6 c	<p>An explanation that links the following six points</p> <p>chlorine solution and potassium bromide solution</p> <p>M1 (solution) turns orange</p> <p>M2 (because) chlorine displaces bromine</p> <p>M3 (so) chlorine is more reactive than bromine</p> <p>bromine solution and potassium iodide solution</p> <p>M4 (solution) turns brown</p> <p>M5 (because) bromine displaces iodine</p> <p>M6 (so) bromine is more reactive than iodine</p>	<p>ACCEPT yellow REJECT brown/red</p> <p>ACCEPT correct word equation, balanced chemical equation or ionic equation.</p> <p>ACCEPT correct word equation, balanced chemical equation or ionic equation</p> <p>If incorrect use of chloride, bromide or iodide in any marking point deduct 1 mark.</p> <p>ALLOW 1 mark for correct order of reactivity given if M3 and M6 not scored</p>	<p>6</p> <p>Total 13</p>

Question number	Answer	Notes	Marks
7 (a) (i)	$\text{Zn (s)} + \text{H}_2\text{SO}_4\text{(aq)} \rightarrow \text{ZnSO}_4\text{(aq)} + \text{H}_2\text{(g)}$	ACCEPT upper case letters	1
(ii)	effervescence/bubbles/fizzing	ACCEPT zinc gets smaller or disappears IGNORE hydrogen / gas produced / given off	1
(b) (i)	An explanation that links the following two points M1 to make sure all of the acid reacts M2 (so that) a pure zinc sulfate solution is obtained/pure zinc sulfate crystals are obtained OWTTE		2
(ii)	M1 filter funnel containing filter paper M2 suitable container to collect filtrate e.g. beaker, conical flask, evaporating basin	M2 dep on a filter funnel in M1	2

Question number	Answer	Notes	Marks
7 (c) (i)	<ul style="list-style-type: none"> • setting out of calculation • evaluation <p>Example calculation</p> <p>M1 $65 + 32 + (4 \times 16) + (7 \times 18)$</p> <p>M2 287</p>	correct answer without working scores 2	2
(ii)	<ul style="list-style-type: none"> • multiply moles by M_r • evaluation <p>Example calculation</p> <p>M1 (mass of $\text{ZnSO}_4 \cdot 7\text{H}_2\text{O}$ =) $287 \times 0.02(00)$</p> <p>M2 (mass of $\text{ZnSO}_4 \cdot 7\text{H}_2\text{O}$ =) 5.74 (g)</p>	correct answer (5.74) without working scores 2	2
(iii)	<p>M1 $4.28 \div 5.74$ OR 0.7456</p> <p>M2 0.7456×100</p> <p>M3 74.6</p>	<p>correct answer to 3 sig figs without working scores 3</p> <p>ALLOW ecf from (ii)</p> <p>ALLOW use of 6g</p> <p>ALLOW any number of sig fig greater than 1</p> <p>ALLOW use of 6g giving answer of 71.3</p> <p>Must be 3 sig figs to score M3</p>	3
			Total 13

Question number	Answer	Notes	Marks
8 (a)	(i) M1 white flame/light M2 white powder/solid (formed)	ALLOW white smoke ALLOW white ash REJECT white precipitate	2
	(ii) magnesium gains oxygen	ACCEPT magnesium loses electrons/ oxidation state (of Mg) increases / goes from / changes from 0 to +2	1
(b)	(i) gives out/releases heat (energy)/thermal energy	IGNORE energy alone	1
	(ii) $2\text{Mg} + \text{CO}_2 \rightarrow 2\text{MgO} + \text{C}$	ALLOW multiples and fractions	1
	(iii) Any one from the fire would keep burning OR the carbon dioxide would not put out the fire OR a large amount of heat/thermal energy would be released	ALLOW Any other sensible suggestion	1

Question number	Answer	Notes	Marks
8 (c)	<p>An explanation giving two linked changes</p> <p>M1 (the student should) lift and replace the lid</p> <p>M2 (to allow) oxygen/air to enter the crucible (to react with the magnesium)</p> <p>AND</p> <p>M3 reheat and reweigh / heat to constant mass</p> <p>M4 to make sure that all the magnesium has reacted</p>		4
			Total 10

Question number	Answer	Notes	Marks
9 (a) (i)	B 5 A is incorrect as 1 is the pH of a strong acid C is incorrect as 7 is the pH of a neutral solution D is incorrect as 9 is the pH of a weak alkali		1
(ii)	D acids are proton donors A is incorrect as alkalis contain OH ⁻ ions not acids B is incorrect as acids do not donate electrons C is incorrect as bases are proton acceptors not acids		1
(b) (i)	(thermal) decomposition		1
(ii)	$\text{PbCO}_3 \rightarrow \text{PbO} + \text{CO}_2$	ALLOW multiples	1

Question number	Answer	Notes	Marks
10 (a) (i)	potassium hydroxide + hydrochloric acid → potassium chloride + water (ii) M1 to mix (the two solutions more thoroughly) M2 (so that) more reactant particles come into contact with each other OWTTE M3 so that the heat energy is given out more quickly OWTTE M4 so that the mixture is the same temperature throughout OWTTE	ALLOW correctly balanced chemical equation ALLOW references to increasing rate of reaction	1 2
(b)	Correct answer with or without working scores 2 <ul style="list-style-type: none"> • setting out of calculation • evaluation Example calculation M1 $\frac{17.8 + 18.4}{2}$ M2 18.1		2

Question number	Answer	Notes	Marks
10 (c) (i)	<p>Mean temperature at start in °C 17.2</p> <p>Temperature at end in °C 22.4</p> <p>Temperature rise in °C 5.2</p>		2
(ii)	<ul style="list-style-type: none"> • calculation of volume/mass of mixture • substitution of values into $Q = mc\Delta T$ • evaluation <p>Example calculation</p> <p>M1 (volume/mass =) 25 + 25 OR 50 (cm³) or (g)</p> <p>M2 ($Q =$) 50 x 4.2 x 5.2</p> <p>M3 ($Q =$) 1092</p>	<p>ALLOW ecf if 25 used in calculation</p> <p>1092 without working scores 3 marks</p>	3
(iii)	<ul style="list-style-type: none"> • division of Q by moles of KOH • conversion of J to kJ • answer with correct sign <p>Example calculation</p> <p>M1 $\frac{1092}{0.02}$ OR 54600</p> <p>M2 conversion from J to kJ OR 54.6(00)</p> <p>M3 ($\Delta H =$) - 54.6 (kJ/mol)</p>	<p>ALLOW ecf from answer to (ii)</p> <p>ALLOW any number of sig figs greater than 1 throughout</p> <p>ACCEPT 1092 or 1100 used in calculation</p> <p>ALLOW ecf from M1</p> <p>Minus sign must be present</p> <p>ALLOW ecf from M2</p> <p>ACCEPT any value between 54.5 and 55</p> <p>M3 dep on division of Q by moles</p>	3

		Correct answer with correct sign and without working scores 3	
		Correct answer without sign or with incorrect sign and without working scores 2.	
			Total 13

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